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Chapter 1 Chemical bonding and chemical

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electrons to give non-hydrogen atoms full

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octets Formal charge = (Valence electrons) - (nonbonding electrons) - ( $\frac{1}{2}$  bonding electrons) Nonpolar when electrons are shared evenly More eneg in top right (More eneg, more attraction towards other atoms) Molecular dipole - Sum of individual dipoles 1.3 Bond length decreases left to ...

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